**Section I: Lab Overview**

The purpose of the lab was to *[insert your objective].* The materials we used for the investigation included: *[summarize your materials].* We performed the investigation by *[summarize in 2 or 3 sentences was you did].* We shared the work by *[summarize how you shared the work in 2 or 3 sentences].*

**Section I: Claims**

Part 1. We think the first reaction was *[endothermic or exothermic]* because our ΔH for this reaction is [*insert the value you calculated in part 1, step 4].*

Part 2. We think the second reaction was *[endothermic or exothermic]* because our ΔH for this reaction is [*insert the value you calculated in part 2, step 4].*

Overall. The overall ΔH for both reactions is [*insert the value you calculated in conclusions, step 3].*

**Section III: Evidence**

Part 1. We gathered our evidence for part 1 in a data table as follows:

|  |  |
| --- | --- |
| Mass of anhydrous pellet (or powder) |  |
| Mass of water (g) |  |
| Initial temperature of water (°C) |  |
| Final temperature of water (°C) |  |
| **Δ**Twater= TFinal - TInitial |   |

Part 2. We gathered our evidence for part 2 in a data table as follows:

|  |  |
| --- | --- |
| Mass of hydrated crystals |  |
| Mass of water (g) |  |
| Initial temperature of water (°C) |  |
| Final temperature of water (°C) |  |
| **Δ**Twater= TFinal - TInitial |   |

**Section IV: Reasoning**

Our evidence [*does / does not]* support our claim because the theoretical change of enthalpy *ΔH* for both reactions combined is -104kJ, and our calculation is *[insert overall ΔH for both reactions shown above from conclusions, step 3].* Our percent error is [*insert the value you calculated in conclusions, step 4].* This makes sense because *[name the relevant scientific law or theory about energy and enthalpy – look in your notes or textbook and summarize what the law/theory says].*

Summary: We summarize our reasoning with the following diagrams for experiment 1 and experiment 2 [*insert your drawings from applications, ~~step 7~~]*.